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3 (Sem-4/CBCS) CHE HC3

2023

CHEMISTRY

(Honours Core)

Paper : CHE-HC-4036

(Physical Chemistry-IV)

Full Marks : 60

Time : Three hours

The figures in the margin indicate full marks for the questions.

1. Answer the following questions : $1 \times 7 = 7$

- (a) What weight of AlF_3 salt be dissolved in 100 ml of solution so as to make the solution containing 1 eq/L ?
- (b) Define equivalent conductance.
- (c) What is cell constant ?
- (d) What is transport number ?

Contd.

(e) Ionic product of water at 25°C is approximately equal to

(i) $1 \times 10^{-7} (\text{mol L}^{-1})^2$

(ii) $2 \times 10^{-14} (\text{mol/L})^2$

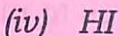
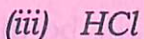
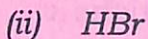
(iii) $1 \times 10^{-14} \text{ mol}^2 \text{ L}^{-2}$

(iv) $1 \times 10^{-7} \text{ mol}^2 \text{ dm}^{-6}$

(Choose the correct answer)

(f) Write two categories of electrochemical cell.

(g) Which of the following hydrogen halides has most polar bond ?



(Choose the correct answer)

2. Answer following questions : $2 \times 4 = 8$

(a) Find the relationship between molar conductance and specific conductance in SI unit.

(b) A perfectly cubical conductivity cell holds 0.94 cm^3 of a solution between its electrodes. Determine its cell constants.

(c) What is relaxation effect ?

(d) Write precisely on potentiometric titration.

3. Answer **any three** questions from the following : $5 \times 3 = 15$

(a) Discuss the Arrhenius theory of electrolytic dissociation. Give evidence in support of the dissociation theory.

$3 + 2 = 5$

(b) Write the principle of conductometric titrations. Discuss the characteristics of curves obtained in the titration of *any two* given below : $1 + (2 + 2) = 5$

(i) HCl vs NaOH

(ii) CH_3COOH vs NaOH

(iii) HCl vs NH_4OH

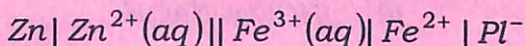
(iv) CH_3COOH vs NH_4OH

(c) (i) What is ionic mobility ? What is the effect of temperature on ionic mobility ? 2

(ii) A potential of 12.0 volts was applied to two electrodes placed 20 cm apart. A dilute solution of NH_4Cl was placed between the electrodes when NH_4^+ is found to cover a distance of 1.6 cm in one hour. What is the mobility of NH_4^+ ion ? 3

(d) (i) Derive a mathematical relation between the electrical energy of reversible galvanic cell and in free energy of the cell reaction. 3

(ii) What is half cell reaction ? Write the half cell reaction of the following cell : 2



(e) Briefly explain Gouy's method for the measurement of magnetic susceptibility.

4. Answer **any three** questions from the following: 10×3=30

(a) (i) How can you measure electrolytic conductance, specific conductance, equivalent conductance and molar conductance ? Write the unit of cell constant (K) in SI unit.

(ii) The resistance of 0.01 M solution of an electrolyte was found to be 210 ohm at 25 °C. Calculate the molar conductance of the solution at 25 °C.

(Given : cell constant = 0.88 cm^{-1})

(iii) Specific conductance of an electrolyte solution decreases with dilution. Explain.

5+3+2=10

(b) (i) State and explain the Kohlrausch's law of independent migration of ions.

- (ii) For the strong electrolytes NaOH , NaCl and BaCl_2 the molar ionic conductance at infinite dilution are 248.1×10^{-4} , 126.5×10^{-4} and $280.0 \times 10^{-4} \text{ S m}^2 \text{ mol}^{-1}$ respectively. Calculate Λ_m° for $\text{Ba}(\text{OH})_2$.
- (iii) Illustrate the application of Kohlrausch's law. $5+2+3=10$
- (c) (i) Illustrate how the solubility product of a sparingly soluble salt can be determined with the help of conductance measurement.
- (ii) What is Ostwald dilution law? Write its verification, importance and limitations. $5+5=10$
- (d) (i) Find the mean ionic activity of a uni-univalent electrolyte.
- (ii) How can you calculate the equilibrium constant of a cell reaction of the type
- $$aA + bB \rightleftharpoons cC + dD ?$$

- (iii) Calculate the equilibrium constant of the cell reaction



occurring in the $Zn - Ag$ cell at $25^\circ C$ when $[Zn^{2+}] = 0.10M$ and $[Ag^+] = 10M$. The EMF of the cell is found to be 1.62 volts.

$$2+5+3=10$$

- (e) (i) State and explain the Nernst equation.
- (ii) Find out whether Zn and Ag would react with dilute H_2SO_4 acid or not.

Given :

$$E_{el}^{\circ} = 0 \text{ for } 2H^+, H_2(g); Pt$$

$$E_{el}^{\circ} = -0.76 V \text{ for } Zn^{2+}; Zn$$

$$E_{el}^{\circ} = +0.80 V \text{ for } Ag^+; Ag$$

$$4+(2 \times 3)=10$$

- (f) (i) How can you apply the dipole moment of a molecule to study its molecular structure ?

- (ii) Find the percentage of ionic character of HCl molecule using SI unit.

Given :

Internuclear distance (r) = 127 pm

Electronic charge = $1.6 \times 10^{-19} \text{ C}$

Actual dipole moment =

$3.44 \times 10^{-30} \text{ coulomb metre.}$

- (iii) How can you distinguish diamagnetic substances and paramagnetic substances depending on the behaviour in a magnetic field?

- (iv) Explain polar and nonpolar covalent bonds.

- (v) Explain the variation of molar polarization with temperature.

$$2+2+2+2=10$$